SNC1P  **Element Standard Notation Practice** Date:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Overall Learning Goal:** We are learning to identify the characteristics of neutrons, protons, and electrons and how to represent atomic structure.

**Success Criteria :**

I can use the information about the protons, neutrons and electrons in an atom to write its standard symbol.

**Example:**

Phosphorus has 15 protons and 16 neutrons.

15

Atomic number= \_\_\_\_\_\_\_\_\_ Atomic number= \_\_\_\_\_\_\_\_\_

15 + 16 = 31

Mass number = \_\_\_\_\_\_\_\_\_\_\_ Mass number = \_\_\_\_\_\_\_\_\_\_\_

Standard Notation: Standard Notation:

31

P

15

Practice:

1. Boron has 5 protons and neutrons. 2) Fluorine has 9 protons and 10 neutrons

Atomic number= \_\_\_\_\_\_\_

Mass number = \_\_\_\_\_\_\_\_\_

Standard Notation:

Atomic number= \_\_\_\_\_\_\_

Mass number = \_\_\_\_\_\_\_\_\_

Standard Notation:

3)Silicon has 14 protons and 14 neutrons. 4) Lithium has 3 protons and 4 neutrons

Atomic number= \_\_\_\_\_\_\_

Mass number = \_\_\_\_\_\_\_\_\_

Standard Notation:

Atomic number= \_\_\_\_\_\_\_

Mass number = \_\_\_\_\_\_\_\_\_

Standard Notation:

Practice Continued:

5)Potassium has 19 protons and 20 neutrons. 6) Argon has 18 protons and 22 neutrons

Atomic number= \_\_\_\_\_\_\_

Mass number = \_\_\_\_\_\_\_\_\_

Standard Notation:

Atomic number= \_\_\_\_\_\_\_

Mass number = \_\_\_\_\_\_\_\_\_

Standard Notation:

7)Helium has 10 protons and 10 neutrons. 8) Sulfur has 16 protons and 16 neutrons.

Atomic number= \_\_\_\_\_\_\_

Mass number = \_\_\_\_\_\_\_\_\_

Standard Notation:

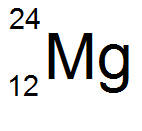
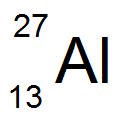
Atomic number= \_\_\_\_\_\_\_

Mass number = \_\_\_\_\_\_\_\_\_

Standard Notation:

Use the Standard Notation for the elements below to determine the number of PROTONS, NEUTRONS and ELECTRONS for each atom assuming the atom is NEUTRAL.

1. Standard Notation for Aluminum: B)Standard Notation for Magnesium:



# protons= \_\_\_\_\_\_\_\_\_\_\_ # protons= \_\_\_\_\_\_\_\_\_\_\_

# neutrons = \_\_\_\_\_\_\_\_\_\_ # neutrons = \_\_\_\_\_\_\_\_\_\_

# electrons= \_\_\_\_\_\_\_\_\_\_ # electrons= \_\_\_\_\_\_\_\_\_\_